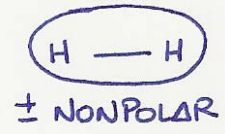
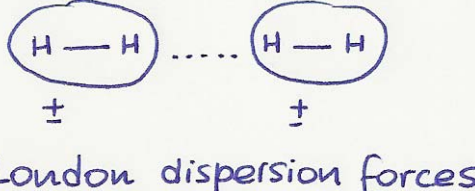


EXERCISE. Complete the information below:				
Formula	Lewis Structure	Shape	Polarity	Intermolecular force
BCl_3		 trigonal planar $\text{Cl}-\text{B}-\text{Cl}$ 120°	 \pm nonpolar	 London dispersion forces
CO_2		 Linear $\text{O}-\text{C}-\text{O}$ 180°	 \pm NONPOLAR	 London dispersion forces
CH_4		 tetrahedral $\text{H}-\text{C}-\text{H}$ 109°	 \pm NonPOLAR	 London dispersion forces
CH_3Cl		 tetrahedral $\text{H}-\text{C}-\text{H}$ 109° $\text{H}-\text{C}-\text{Cl}$	 POLAR	 Dipole-dipole forces
NH_3		 trigonal pyramidal $\text{H}-\text{N}-\text{H}$ 109°	 $\delta+$ POLAR	 hydrogen bonding

EXERCISE. Complete the information below:				
Formula	Lewis Structure	Shape	Polarity	Intermolecular force
PH ₃		trigonal pyramidal H-P-H 109°	± NONPOLAR	± London dispersion forces
H ₂ O		V-shaped H-O-H 109°	δ- POLAR	δ- δ+ hydrogen bonding
SO ₂		V-shaped O-S-O 120°	δ- POLAR	δ- δ+ Dipole-dipole forces
SO ₃		trigonal planar O-S-O 120°	± NONPOLAR	± London dispersion forces
HCl		-----	δ+ POLAR	δ+ δ- Dipole-dipole forces

EXERCISE. Complete the information below:				
Formula	Lewis Structure	Shape	Polarity	Intermolecular force
HF		-----		
H ₂ CO				
HCN				
H ₂ S				
Cl ₂		-----		

EXERCISE. Complete the information below:				
Formula	Lewis Structure	Shape	Polarity	Intermolecular force
H ₂	H — H	-----		
Cl ₂ O	