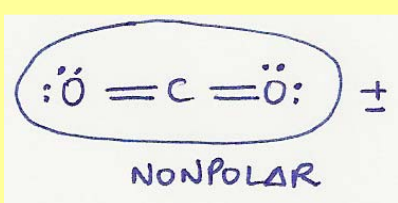
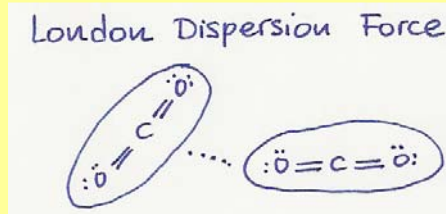
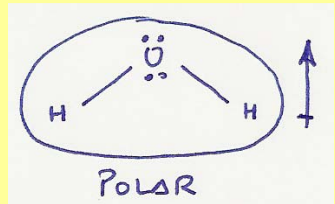
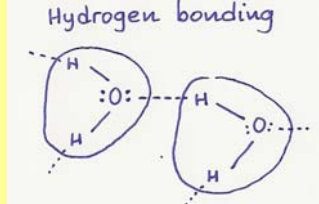
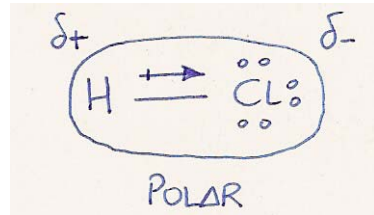
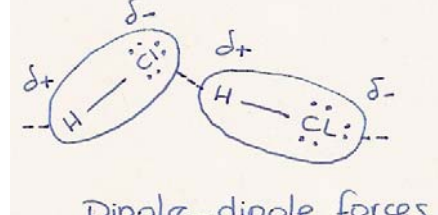
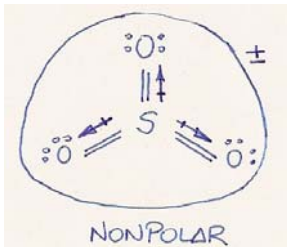
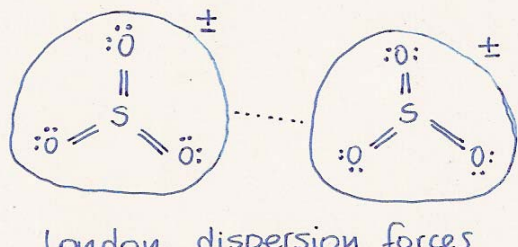
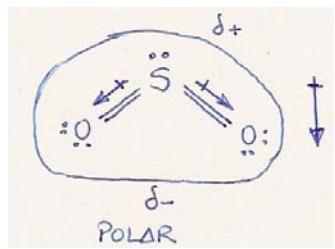
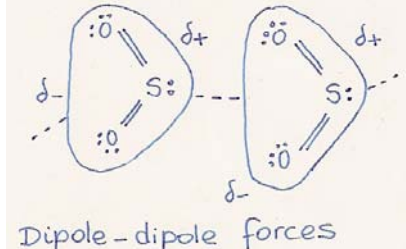
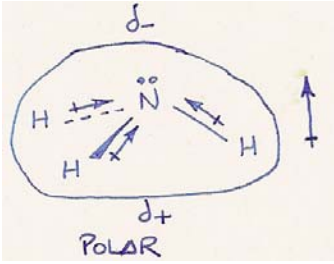
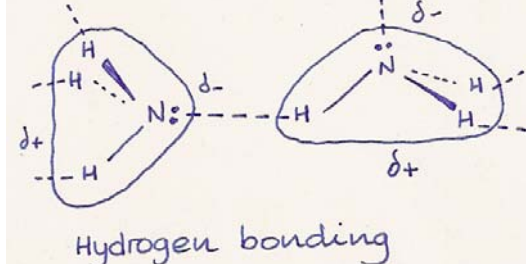
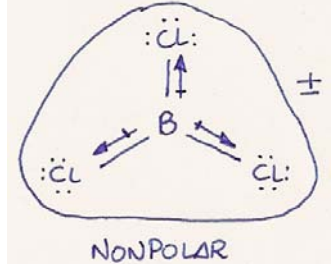
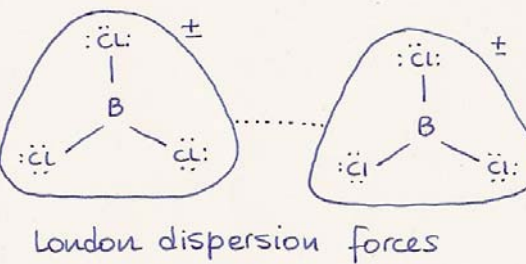


Topic:	Intermolecular Forces
Objective:	FK_03_06
<p>Given a molecule and the electronegativity values the student must be capable of doing the following:</p> <ul style="list-style-type: none"> determine the polarity of each bond determine the polarity of the molecule (polar / nonpolar) determine its behavior between positive and negative plates 	

EXERCISE. Determine the type of intermolecular force that exists in the following examples.

Formula	Molecular Polarity	Intermolecular Force
CO ₂	 <p>NONPOLAR</p>	<p>London Dispersion Force</p> 
H ₂ O	 <p>POLAR</p>	<p>Hydrogen bonding</p> 
HCl	 <p>POLAR</p>	 <p>Dipole-dipole forces</p>
SO ₃	 <p>NONPOLAR</p>	 <p>London dispersion forces</p>
SO ₂	 <p>POLAR</p>	 <p>Dipole-dipole forces</p>

<p>NH₃</p>	 <p>POLAR</p>	 <p>Hydrogen bonding</p>
<p>BCl₃</p>	 <p>NONPOLAR</p>	 <p>London dispersion forces</p>
<p>Electronegativities: C=2.5; O=3.5; H=2.1; Cl=3.0; S=2.5; N=3.0; B=2.0</p>		