

EXERCISE: determine the type of ions formed and the formula of the compound in the following cases

K + O	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>
Na + S	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>
S + Ca	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>
Mg + O	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>
Cl + Mg	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>
Na + N	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p> <p> $M X$ $M X_2$ $M X_3$ $M_2 X$ $M_3 X$ $M_2 X_3$ </p>

Cl + Na	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>
K + F	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>
Br + K	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>
Al + F	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>
S + Al	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>
Mg + O	<p style="text-align: center;">Ions</p>	<p style="text-align: center;">Formula</p>