Topic:	**Molecular Shape:** **VSEPR Theory (Valence-Shell Electron-Pair Repulsion)**
Objective:	**FK_03_04**

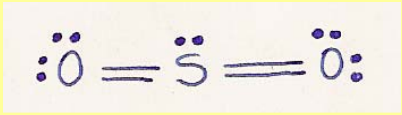
Given a molecule
the student must be capable of doing the following:

- **determine its Lewis structure**
- **determine its shape and bond angles**

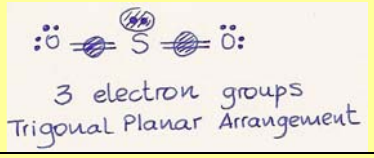
Determine the Lewis structure and the molecular shape of the following compounds

SO_2

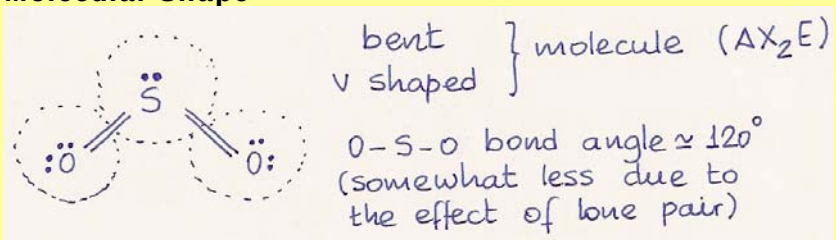
Lewis



Arrangement of Electron Pairs

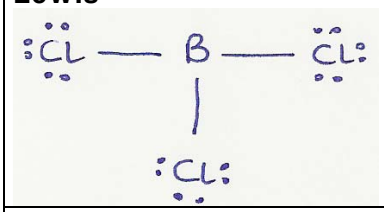


Molecular Shape

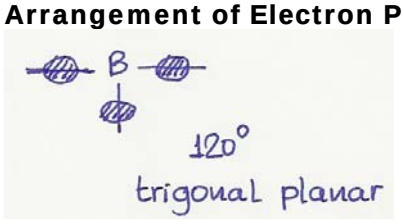


BCl_3

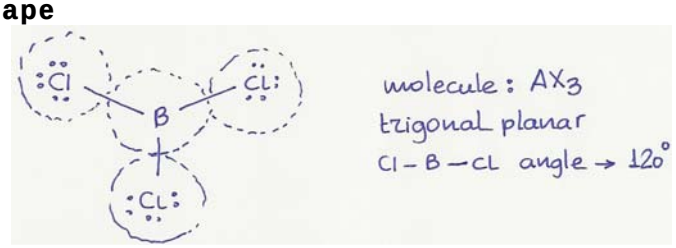
Lewis



Arrangement of Electron Pairs

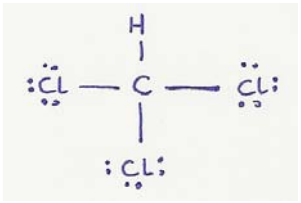


Molecular Shape

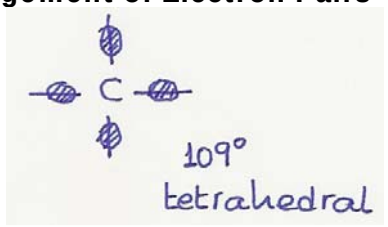


$CHCl_3$

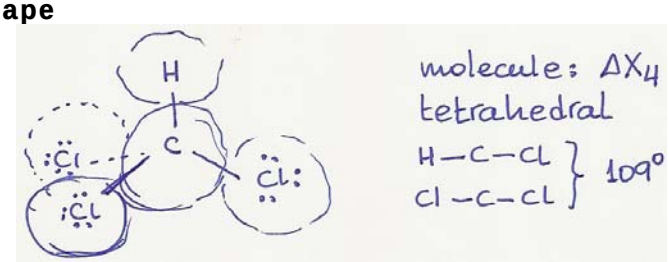
Lewis



Arrangement of Electron Pairs



Molecular Shape



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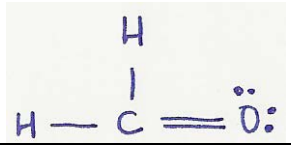
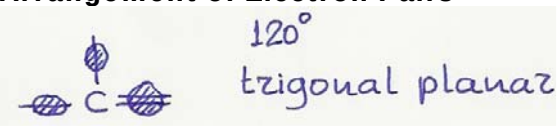
  </td><td>Arrangement of Electron Pairs
  </td></tr>
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 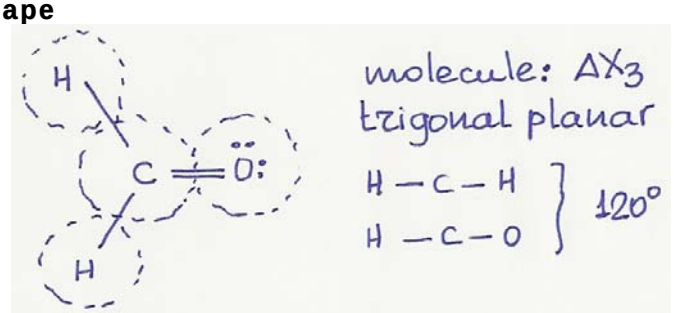 </td></tr>
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 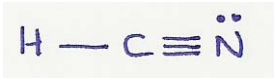 </td><td>Arrangement of Electron Pairs
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 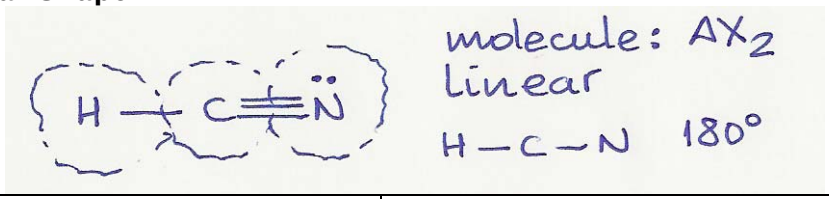 </td></tr>
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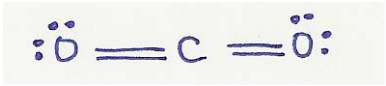
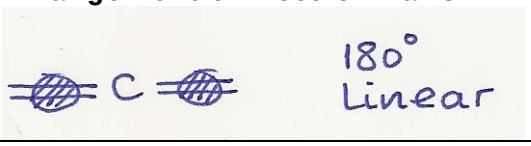
  </td><td>Arrangement of Electron Pairs
  </td></tr>
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 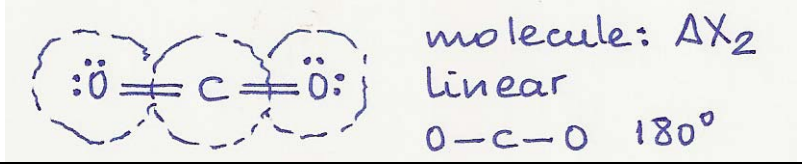 </td></tr>
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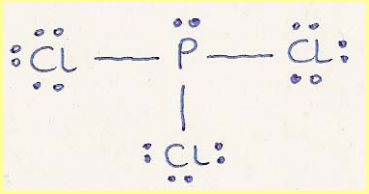
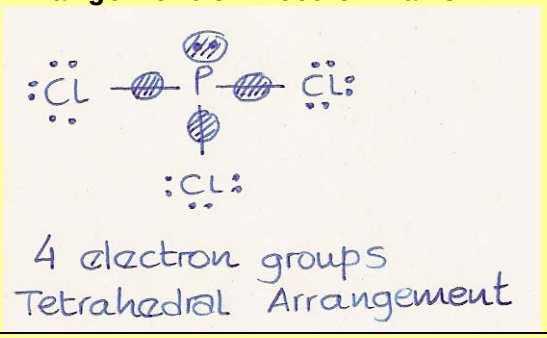
PCl_3

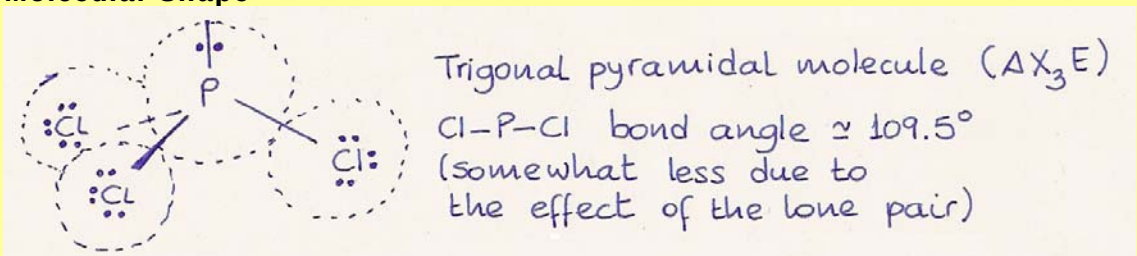
Lewis



Arrangement of Electron Pairs



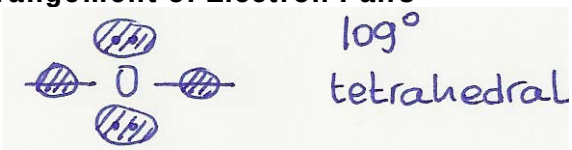
Molecular Shape



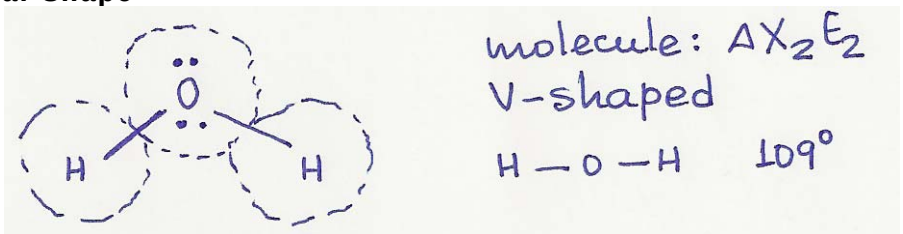
H_2O

Lewis

Arrangement of Electron Pairs

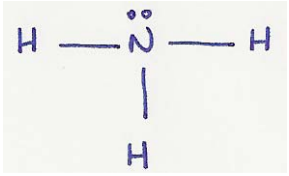


Molecular Shape

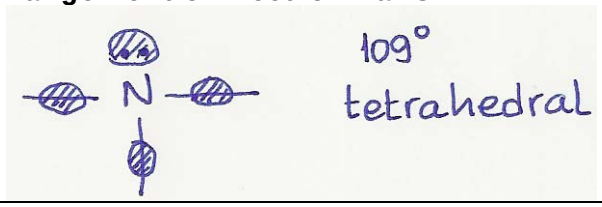


NH_3

Lewis



Arrangement of Electron Pairs



Molecular Shape