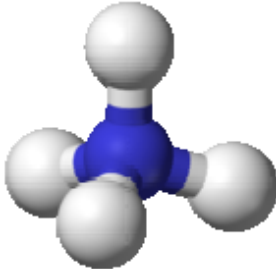
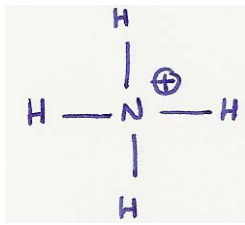
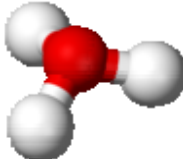
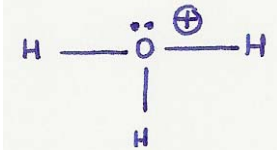
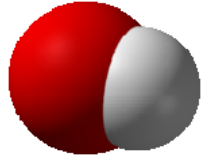

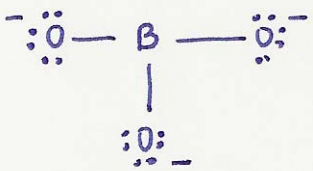
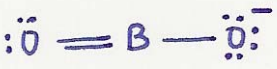
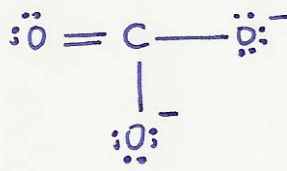
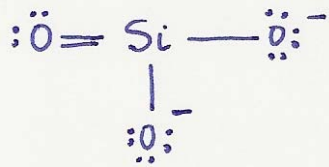
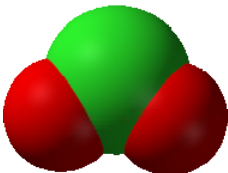
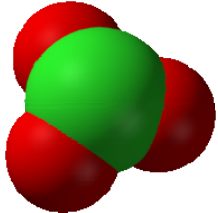
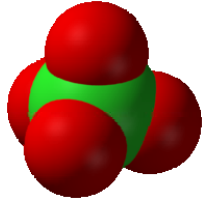


<b>Topic:</b>	<b>Polyatomic ions</b>
<b>Objective:</b>	<b>FK_03_05</b>
<b>Given the name/structure/formula of an polyatomic ion the student must be capable of doing the following:</b> <ul style="list-style-type: none"> <li>• <b>give the formula/structure/name of the ion</b></li> </ul>	

<b>Some polyatomic ions</b>			
<b>Formula</b>	<b>Name</b>	<b>Picture</b>	<b>Lewis</b>
$\text{NH}_4^+$	Ammonium		
$\text{H}_3\text{O}^+$	Hydronium		
$\text{OH}^-$	Hydroxide		
$\text{BO}_3^{3-}$	Ortoborate		
$\text{BO}_2^-$	Metaborate		
$\text{CO}_3^{2-}$	Carbonate		
$\text{SiO}_3^{2-}$	Metasilicate		

$\text{SiO}_4^{4-}$	Ortosilicate		
$\text{NO}_2^-$	Nitrite		
$\text{NO}_3^-$	Nitrate		
$\text{PO}_4^{3-}$	Phosphate		
$\text{PO}_3^{3-}$	Phosphite		
$\text{SO}_4^{2-}$	Sulfate		
$\text{SO}_3^{2-}$	Sulfite		
$\text{ClO}^-$	Hypochlorite		

$\text{ClO}_2^-$	Chlorite		$\text{:}\ddot{\text{O}}=\ddot{\text{Cl}}-\ddot{\text{O}}:^-$
$\text{ClO}_3^-$	Chlorate		$\begin{array}{c} \text{:}\ddot{\text{O}}=\ddot{\text{Cl}}-\ddot{\text{O}}:^- \\    \\ \text{:}\ddot{\text{O}}: \end{array}$
$\text{ClO}_4^-$	Perchlorate		$\begin{array}{c} \text{:}\ddot{\text{O}}: \\    \\ \text{:}\ddot{\text{O}}=\text{Cl}-\ddot{\text{O}}:^- \\    \\ \text{:}\ddot{\text{O}}: \end{array}$