

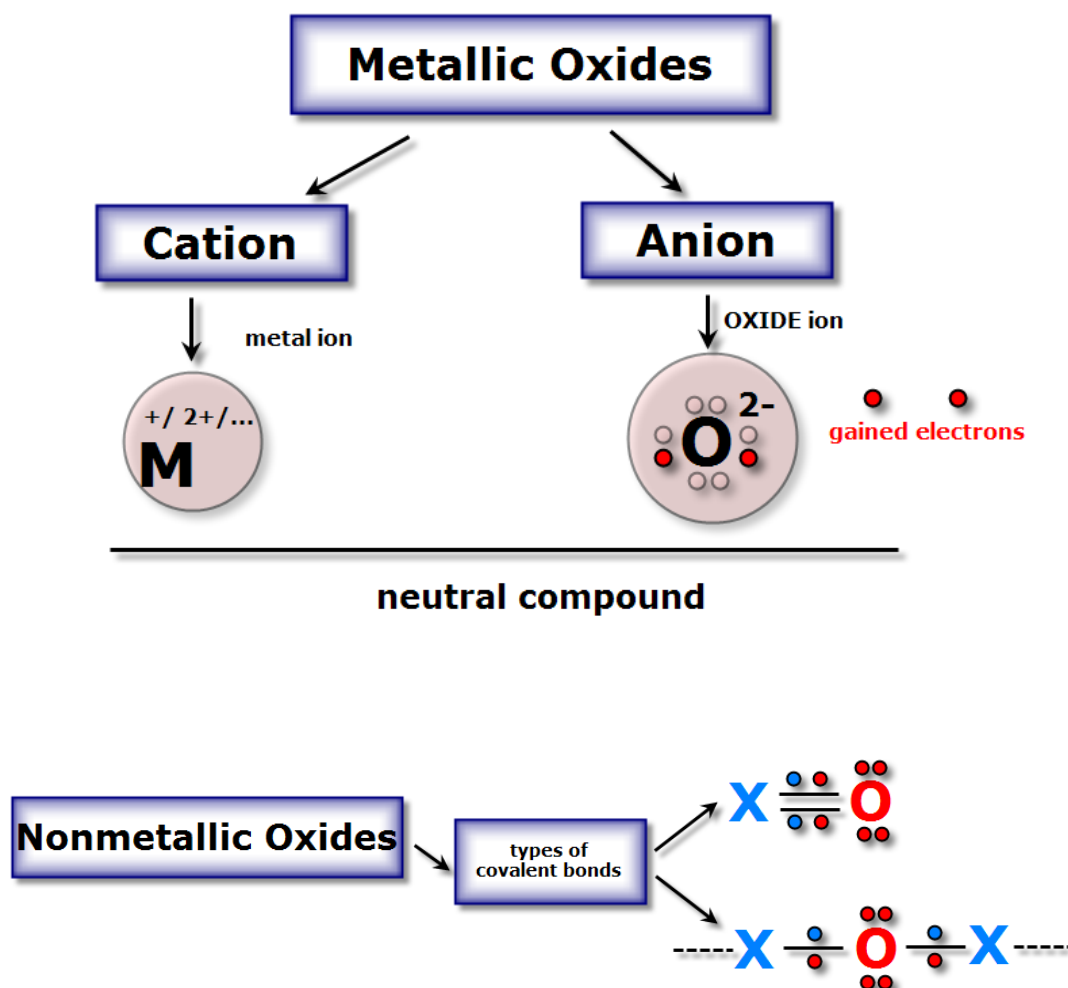
Topic:	OXIDES
Objective:	FK_04_01
Given some compounds and elements the student must be capable of doing the following: <ul style="list-style-type: none"> • write the formulas • write the names • write the Lewis structures 	

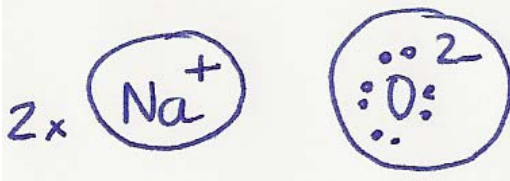
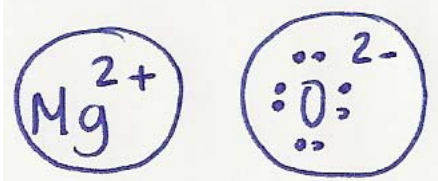
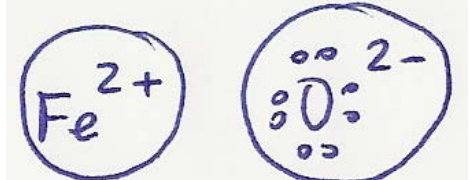
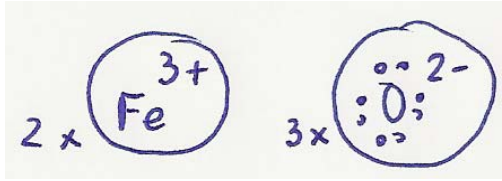
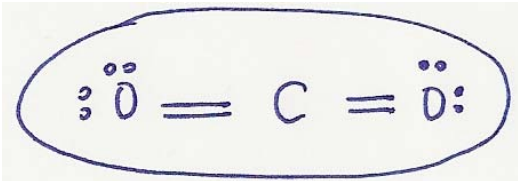
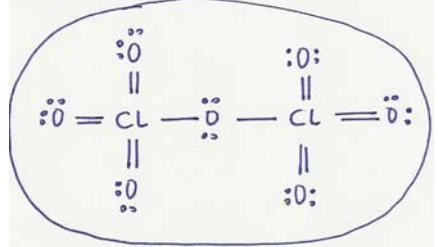
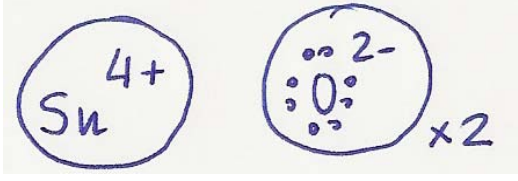
Definition and types

An oxide is a chemical compound containing at least one oxygen atom and other elements (metals or nonmetals).

There are two types of oxides:

- $M + O \dots$ **metallic oxides**. They are ionic compounds.
- $X + O \dots$ **nonmetallic oxides**. These compounds are covalent.



Examples		
Formula	Structure	Name
Na_2O		sodium oxide
MgO		magnesium oxide
FeO		iron(II) oxide
Fe_2O_3		iron(III) oxide
CO_2		carbon dioxide carbon(IV) oxide
Cl_2O_7		Chlorine(VII) oxide Dichlorine heptoxide
SnO_2		tin(IV) oxide tin dioxide

Exercises

Formula	Structure	Name
K_2O		
		calcium oxide
PbO		
		aluminum oxide
CO		
		sulfur trioxide
SO_2		