



# Periodic Table: exercises

Identify (indicating the number) the following elements in the periodic table below

- the element with the symbol "O"
- the element with the symbol "F"
- the element with the symbol "Na"
- the element with the symbol "He"
- the element with the symbol "P"
- the element of the third period that forms a monpositive ion
- the element of the second period that forms a mononegative ion
- the element of the third period that forms a two positive ion
- the element of the second period that forms a two negative ion
- the element with the electron configuration  $1s^2 \dots 3s^2 3p^3$
- the element with the electron configuration  $1s^2 \dots 4s^2$

1	
3	4
11	12
19	20

					2
5	6	7	8	9	10
13	14	15	16	17	18
				21	

## Solutions

- a) 8    b) 9    c) 11    d) 2    e) 15  
f) 11    g) 9    h) 12    i) 8    j) 15    k) 20

# Periodic Table: exercises

Identify (indicating the number) the following elements in the periodic table below

- the element with the symbol "S"
- the element with the symbol "F"
- the element with the symbol "K"
- the element with the symbol "Al"
- the element with the symbol "C"
- the element of the third period that forms a three positive ion
- the element of the second period that forms a mononegative ion
- the element of the third period that forms a monpositive ion
- the element of the second period that forms a two negative ion
- the element with the electron configuration  $1s^2 \dots 4s^2 4p^5$
- the element with the electron configuration  $1s^2 \dots 3s^2$

1						2	
3	4						
11	12						
19	20						
		5	6	7	8	9	10
		13	14	15	16	17	18
						21	

# Periodic Table: exercises

Identify (indicating the number) the following elements in the periodic table below

- the element with the symbol "S"
- the element with the symbol "F"
- the element with the symbol "K"
- the element with the symbol "Al"
- the element with the symbol "C"
- the element of the third period that forms a three positive ion
- the element of the second period that forms a mononegative ion
- the element of the third period that forms a monopositive ion
- the element of the second period that forms a two negative ion
- the element with the electron configuration  $1s^2 \dots 4s^2 4p^5$
- the element with the electron configuration  $1s^2 \dots 3s^2$

1	
3	4
11	12
19	20

					2
5	6	7	8	9	10
13	14	15	16	17	18
				21	

## Solutions

- |       |      |       |       |       |       |
|-------|------|-------|-------|-------|-------|
| a) 16 | b) 9 | c) 19 | d) 13 | e) 6  |       |
| f) 13 | g) 9 | h) 11 | i) 8  | j) 21 | k) 12 |