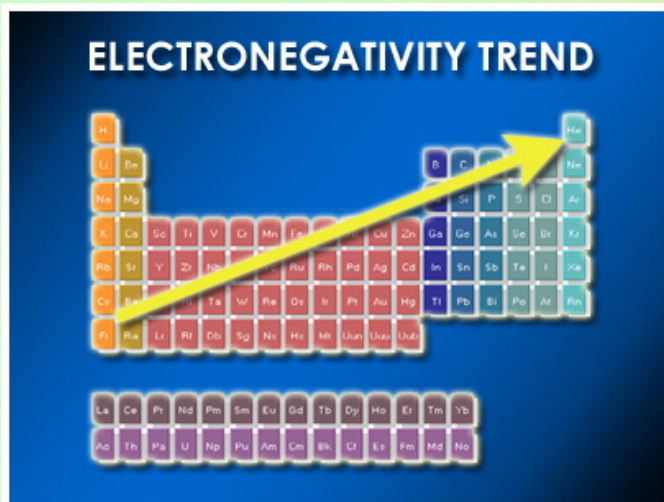


# Electronegativity



<http://www.chem.ubc.ca/courseware/121/tutorials/exp7A/images/fig4.jpg>

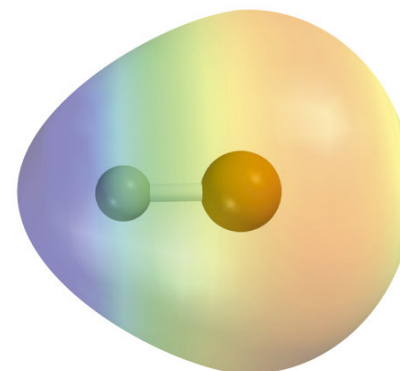
## Definition

Electronegativity is defined as the ability of an atom in a molecule to attract electrons to itself.



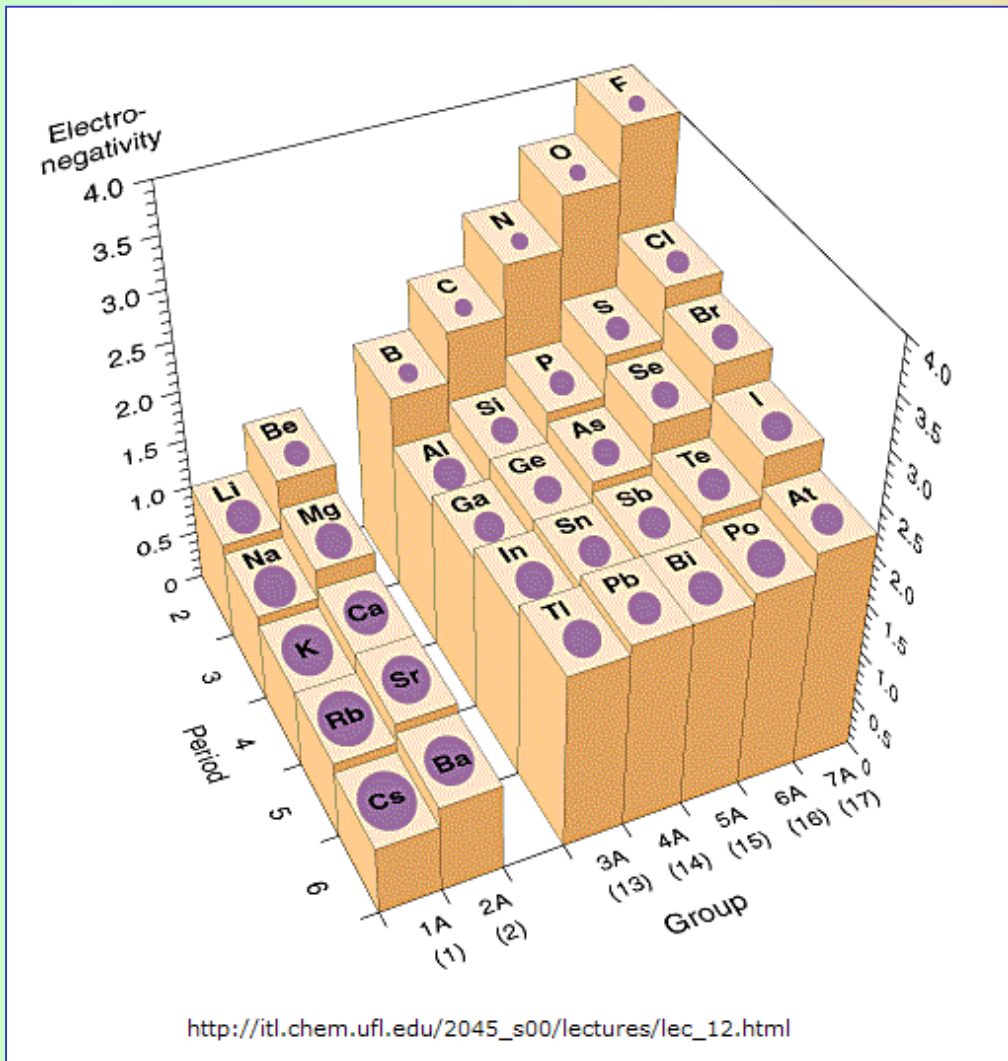
A polar covalent bond.

The bonding electrons are attracted more strongly by Cl than by H.





# Electronegativity



## Trends

In this picture we see the opposite trends of atomic size and electronegativity:

the larger atoms have smaller value of electronegativity.