

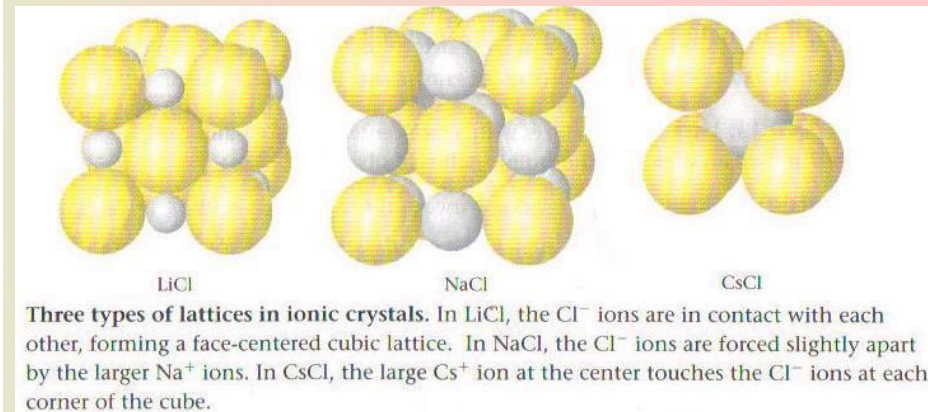
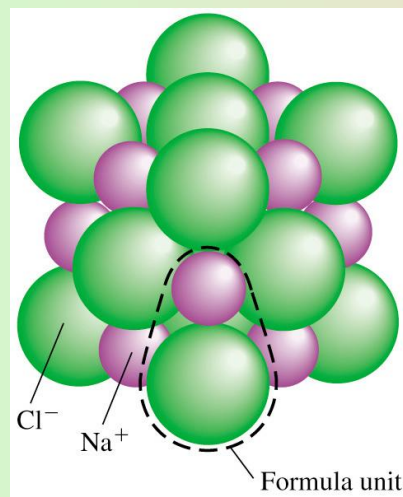
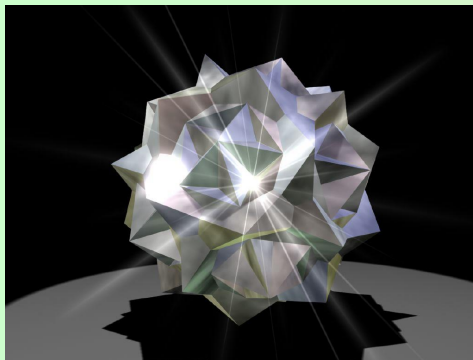
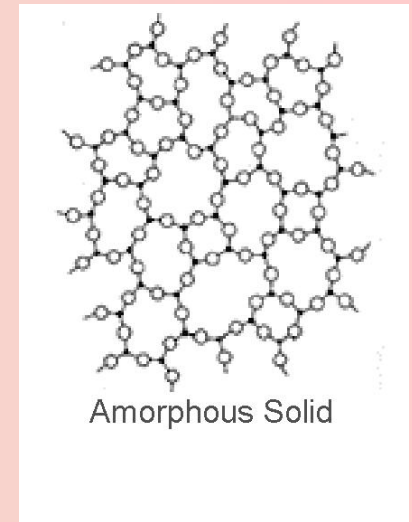
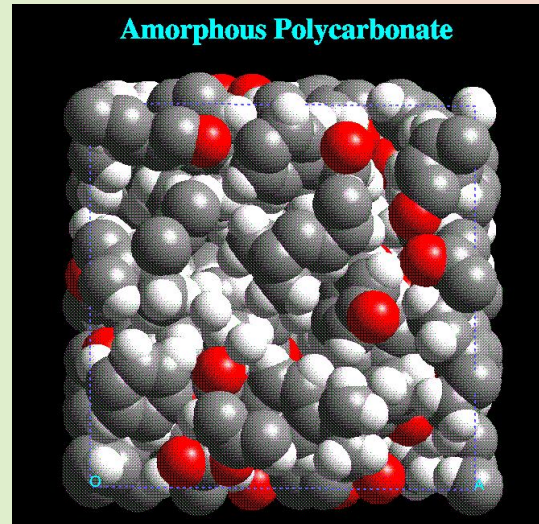
# Kristal egiturak

## Kristal egiturak

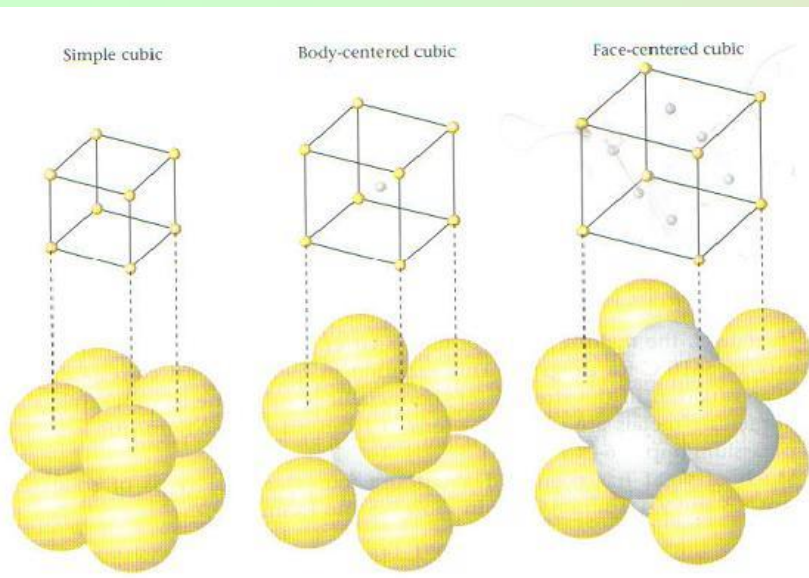
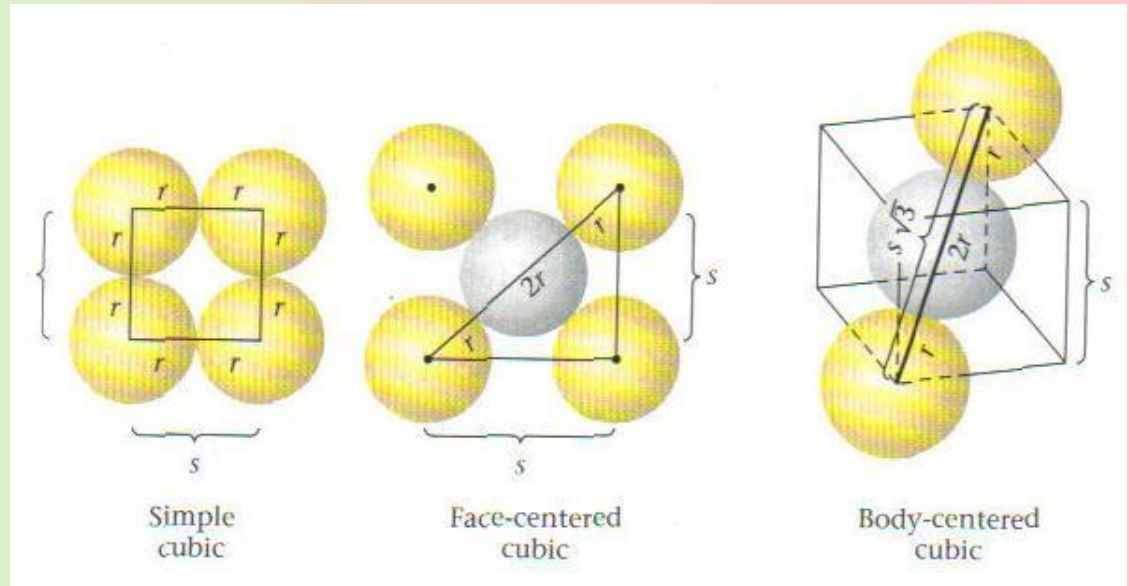
Solidoak era hauetan aurkez daitezke:

egoera amorfoan; partikulek inolako erregulartasunik gabe

kristal egituran; molekulak, atomoak edo ioiak antolatuta forma geometikoren baten arabera, kristal edo sarea osatuz.



# Kristal egiturak



## Properties of Cubic Unit Cells

	Simple	BCC	FCC
Number of atoms per unit cell	1	2	4
Relation between side of cell, $s$ , and atomic radius, $r$	$2r = s$	$4r = s\sqrt{3}$	$4r = s\sqrt{2}$
% of empty space	47.6	32.0	26.0

Helbidea:

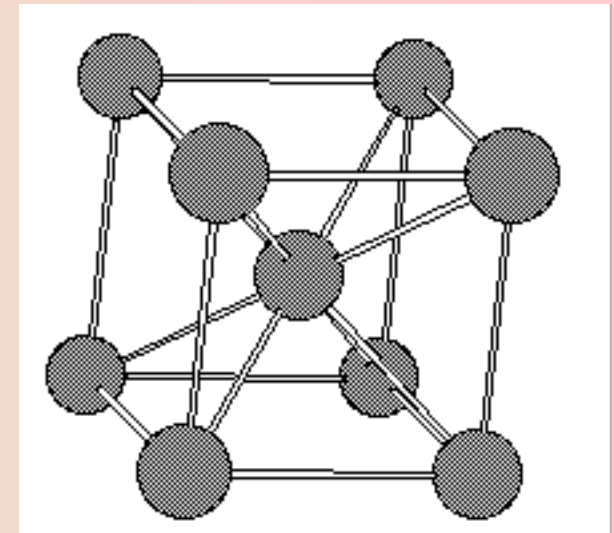
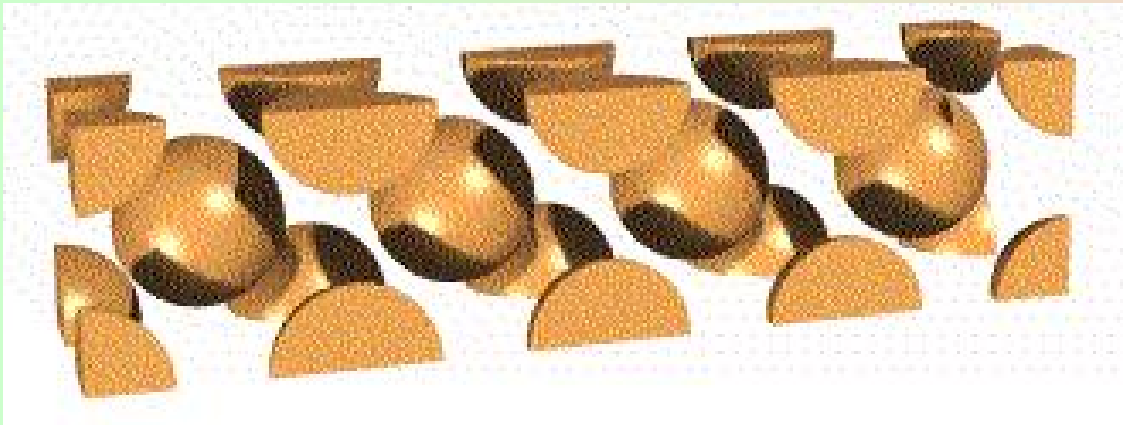
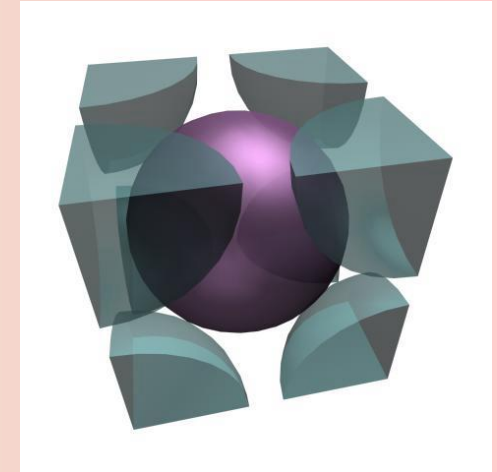
<http://amityregion5.org:2003/chemistry/view.php?top=2&sub=2>

# Kristal egiturak

**Kubiko zentratua (b.c.c. – body-centered cubic):**

atomoak kuboaren erdigunean eta erpinetan kokatzen dira.

Adibidez: Na, K, Fe.

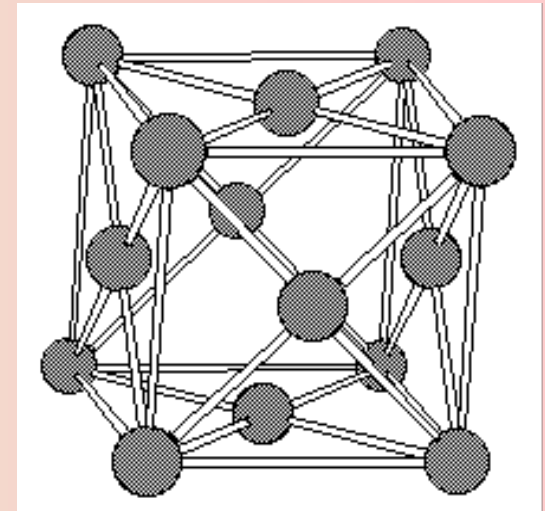
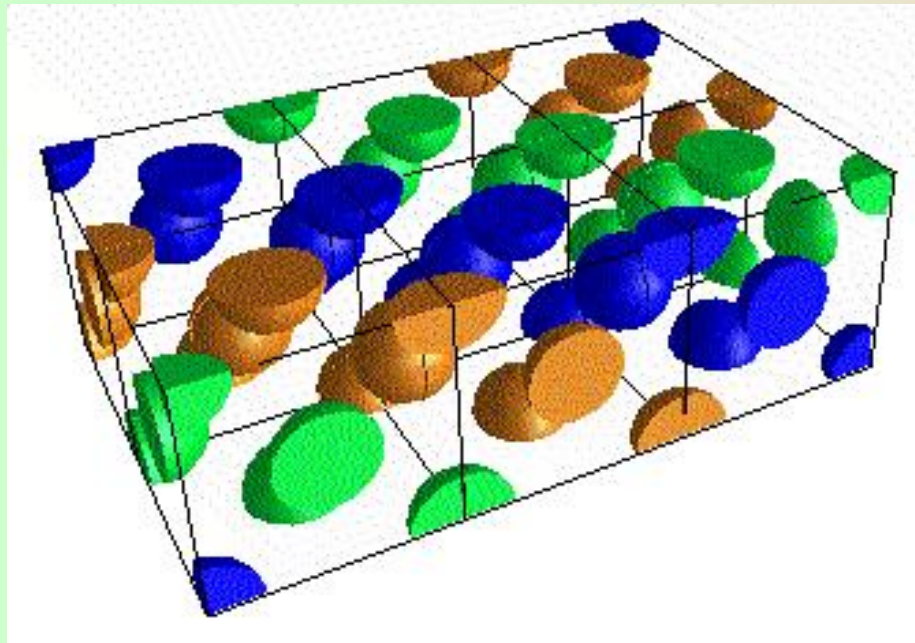
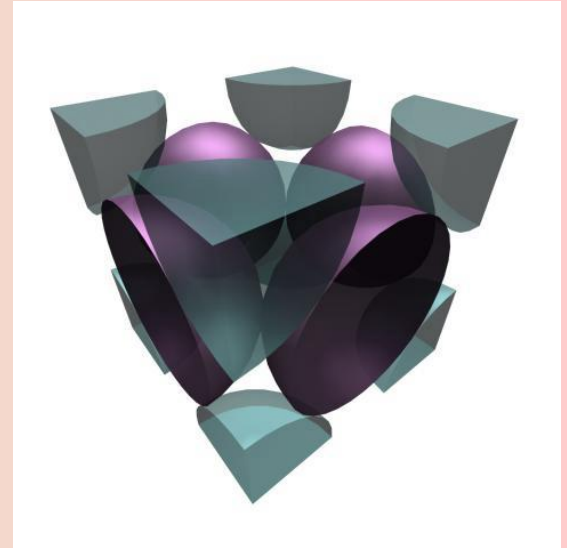


# Kristal egiturak

**Aurpegietan zentratutako kubikoa**  
(f.c.c. – face-centered cubic):

atomoak kuboaren erpinetan eta aurpegien erdigunean kokatzen dira.

Adibidez: Ca, Al, Fe, Cu.



# Kristal egiturak

**Hexagonal trinkoa**  
(h.c.p. – hexagonal closest packing):

atomoak prismaren erpinetan eta oinarrien erdigunean kokatzen dira.

Adibidez: Mg.

