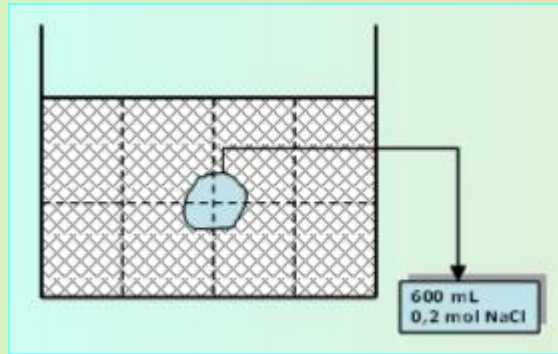
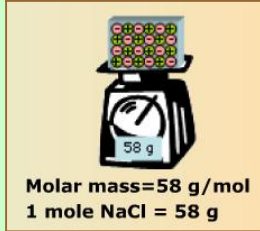


Disoluzioak: kalkuluak



Disoluzio baten bolumena 8 L-koa da eta 600 mL-tan 0,2 mol solutu daude.

balio totalak

m =	g
n =	mol
V =	8 L

propor-tzioak

c =	g/L
	M

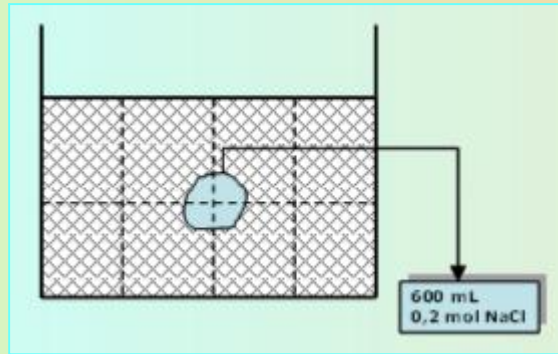
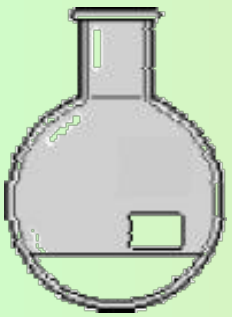
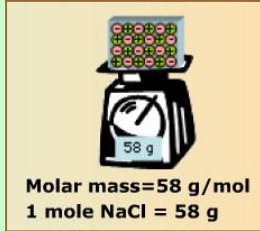
1

1. Datuetatik kontzentrazio molarra kalkulatu dugu.

Laginak eta disoluzio osoak kontzentrazio berdina dute.

$$c = \frac{0,2 \text{ mol}}{0,6 \text{ L dis}} = 0,33 \frac{\text{mol NaCl}}{\text{L dis}}$$

Disoluzioak: kalkuluak



Disoluzio baten bolumena 8 L-koa da eta 600 mL-tan 0,2 mol solutu daude.

balio totalak

m =	g
n =	mol
V =	8 L

propor-tzioak

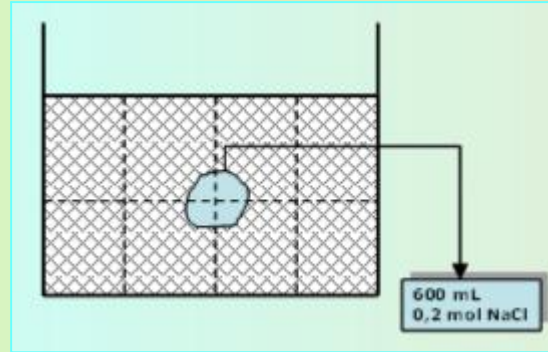
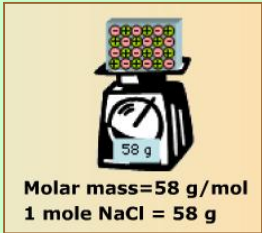
c =	g/L
	0,33 M

2

2. Kontzentratziotik abiatuz, eta bolumena kontuan harturik, solutuaren mol-kopurua kalkulatuko dugu.

$$n = 0,33 \frac{\text{mol NaCl}}{\text{L dis}} * 8 \text{ L} = 2,64 \text{ mol NaCl}$$

Disoluzioak: kalkuluak



Disoluzio baten bolumena 8 L-koa da eta 600 mL-tan 0,2 mol solutu daude.

balio totalak

m =
n =
V =

g
2,64 mol
8 L

3

propor-
tzioak

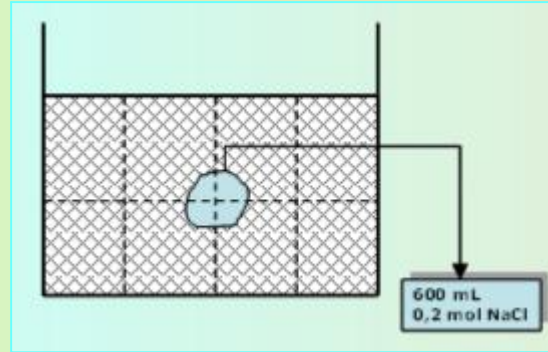
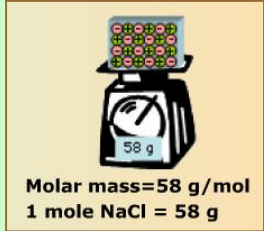
c =

g/L
0,33 M

3. Mol-kopurua jakinik, solutuaren masa kalkulatuko dugu.

$$m = 2,64 \text{ mol NaCl} * \frac{58 \text{ g}}{1 \text{ mol}} = 153,1 \text{ g NaCl}$$

Disoluzioak: kalkuluak



Disoluzio baten bolumena 8 L-koa da eta 600 mL-tan 0,2 mol solutu daude.

balio totalak

$$\begin{array}{l} m = 153,1 \text{ g} \\ n = 2,64 \text{ mol} \\ V = 8 \text{ L} \end{array}$$

4

propor-
tzioak

$$c = \begin{array}{l} 19,1 \text{ g/L} \\ 0,33 \text{ M} \end{array}$$

4. Solutuaren masa eta bolumena jakinik, kontzentrazioa g/L-tan kalkulatu dugu.

$$c = \frac{153,1 \text{ g NaCl}}{8 \text{ L}} = 19,1 \frac{\text{g NaCl}}{\text{L}}$$