


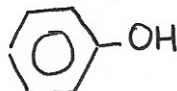
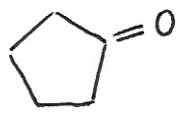
ERREPASO-ARIKETAK #02

Izena

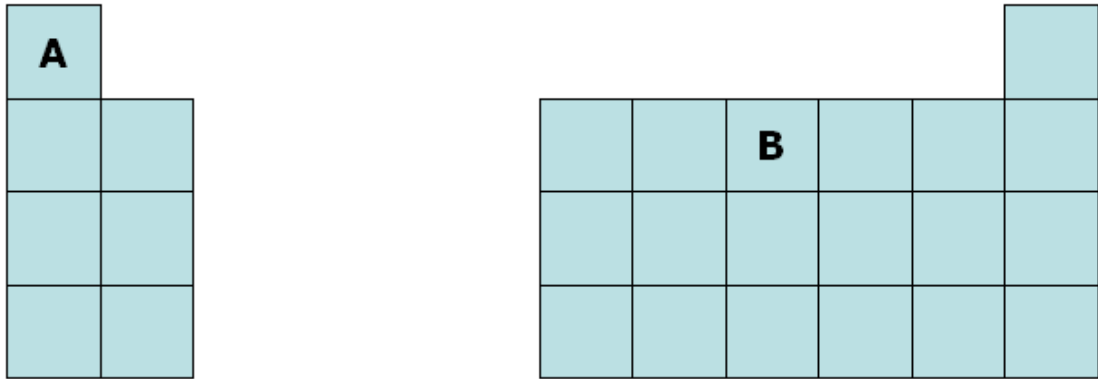
Kurtsoa

1

FORMULAZIOA:
Osatu ondoko taula

Izenak	Formulak
Propeno	$\text{CH}_2 = \text{CH} - \text{CH}_3$
Metil-2-buteno	$\begin{array}{c} \text{CH}_3 - \text{CH} = \text{CH} - \text{CH}_3 \\ \\ \text{CH}_3 \end{array}$
2-butanol	$\begin{array}{c} \text{CH}_3 - \text{CH}_2 - \text{CH} - \text{CH}_3 \\ \\ \text{OH} \end{array}$
Azido azetiko Azido etanoiko	$\begin{array}{c} \text{CH}_3 - \text{C} - \text{OH} \\ \\ \text{O} \end{array}$
Bentzeno	$\begin{array}{c} \text{CH} = \text{CH} \\ / \quad \backslash \\ \text{CH} \quad \quad \text{CH} \\ \backslash \quad / \\ \text{CH} = \text{CH} \end{array}$ edo 
Fenol	
Propanal	$\text{CH}_3 - \text{CH}_2 - \text{CHO}$
Ziklopentanona	

2 TAULA PERIODIKOA eta LOTURAK
Galderak erantzun



1. Adierazi "A" eta "B" elementuen ikurrak eta izenak

A = H / hidrogenoa

B = N / nitrogenoa

2. Eman atomo neutroen konfigurazio elektronikoak kutxa-diagrama formatuan

"A" edo H ... $\boxed{\uparrow}$
1s

"B" edo N ... $\boxed{\uparrow\downarrow}$ $\boxed{\uparrow\downarrow}$ $\boxed{\uparrow}$ $\boxed{\uparrow}$ $\boxed{\uparrow}$
1s 2s 2p

3. Adierazi zein ioi emateko joera izango duten

H ... H^+ eta \ddot{H}^-

N ... $:\ddot{N}:^{3-}$

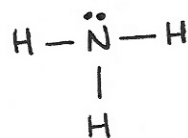
4. Alderatu erradio atomikoak

Ezin da alderatu

5. Alderatu elektronegativotasuna

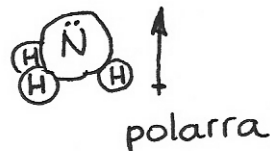
Ezin da alderatu

6. Eman beraien artean osatzen duten konposatuaren formula, izena, molekularen geometria, molekularen polaritatea eta indar intermolekularrak



NH₃
amoniako

piramidala
109°
inguru



indar intermolekularrak:
H zubizko loturak

Ontzi batean 2,5 L azido nitriko (HNO_3) 10 M daukagu. Kalkulatu

1. Zenbat mol azido dauden

$$n(\text{HNO}_3) = 10 \frac{\text{mol}}{\text{L}} \times 2,5 \text{ L} = 25 \text{ mol azido}$$

2. Zenbat gramo azido dauden

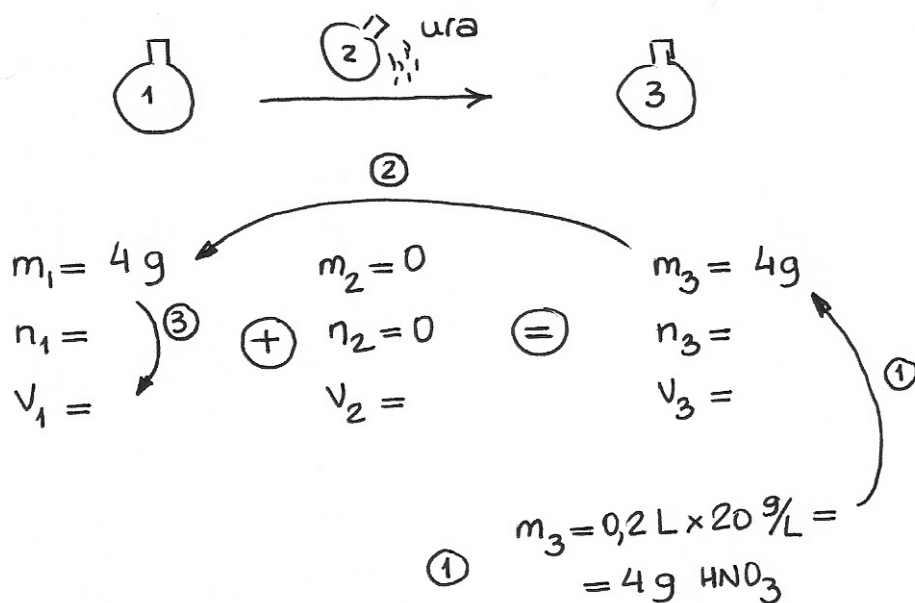
$$M_m(\text{HNO}_3) = 1 + 14 + (3 \times 16) = 63 \text{ g/mol}$$

$$m(\text{HNO}_3) = 25 \text{ mol} \times \frac{63 \text{ g}}{1 \text{ mol}} = 1575 \text{ g azido}$$

3. Kalkulatu zenbat mL hartu beharko diren disoluzio horretatik 200 mL HNO_3 20 g/L lortzeko (diluzioz)

DATUAK

Masa atomikoak: H=1; N=14; O=16



$$\textcircled{2} \quad m_1 + m_2 = m_3$$

$$\begin{array}{c} \xrightarrow{m_2=0} \\ m_1 = 4 \text{ g} \end{array}$$

$$\textcircled{3} \quad v_1 = \frac{2,5 \text{ L}}{1575 \text{ g azido}} \times 4 \text{ g azido} = 0,0063 \text{ L} = \boxed{6,3 \text{ mL}}$$