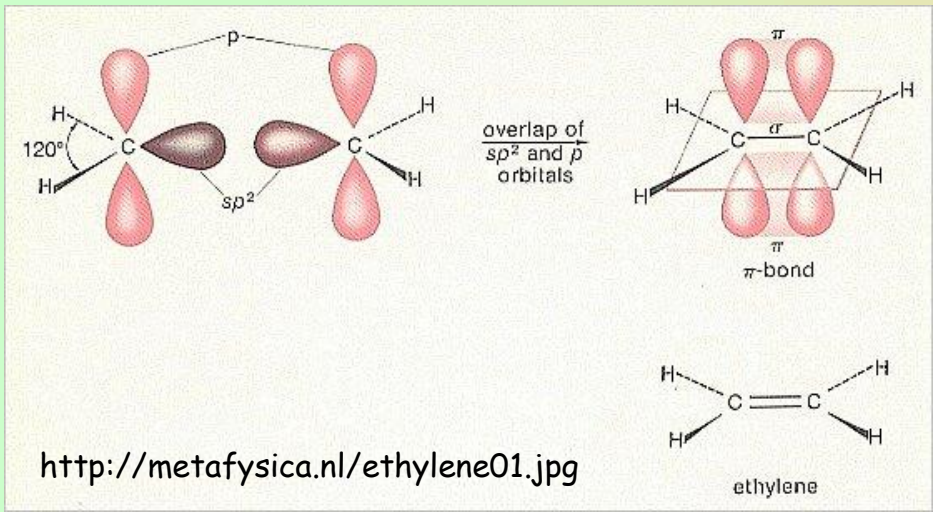


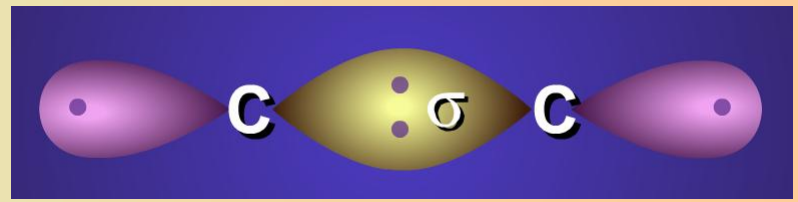
Sigma eta Pi Loturak



Sarrera

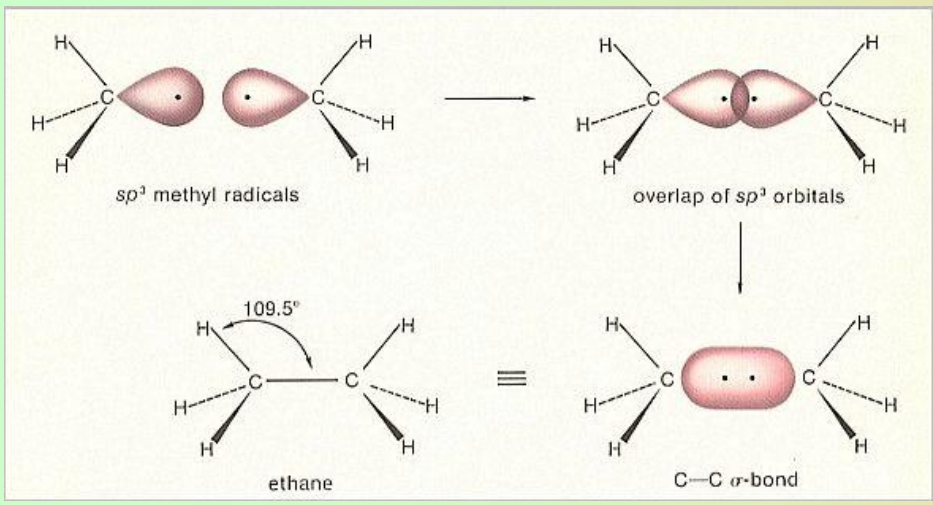
Lotura kobalentea ematen denean orbitalak elkarsartzeko bi era daude:

- Aurrez aurre: (σ) lotura



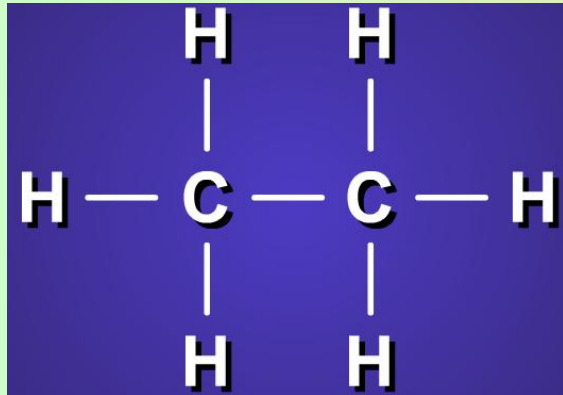
- Alboka: pi (π) lotura

This diagram shows two carbon atoms (C) with their sp^3 hybrid orbitals overlapping end-to-end to form a σ bond. Simultaneously, their unhybridized p orbitals overlap side-by-side to form two π bonds. Labels indicate 'Two lobes of one π bond' and 'Two lobes of one π bond'.



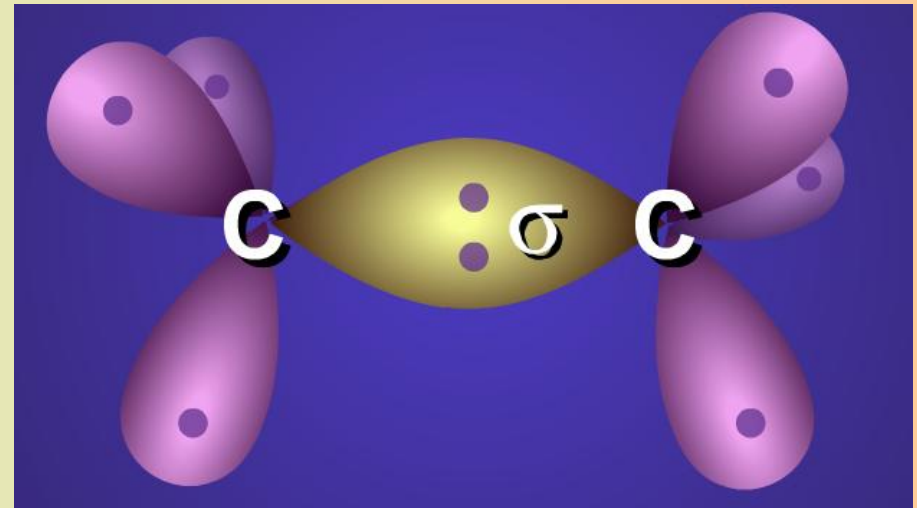
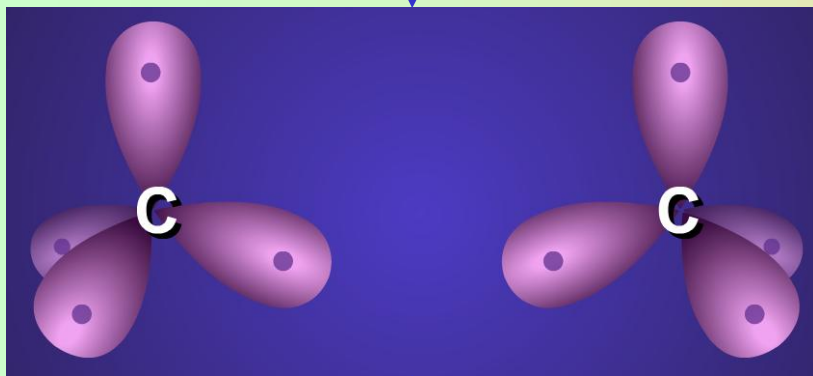
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Sigma eta Pi Loturak

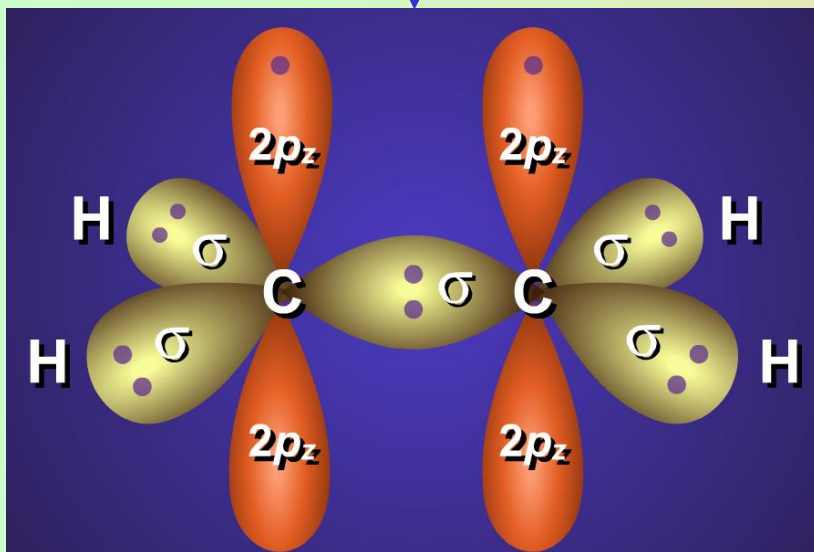
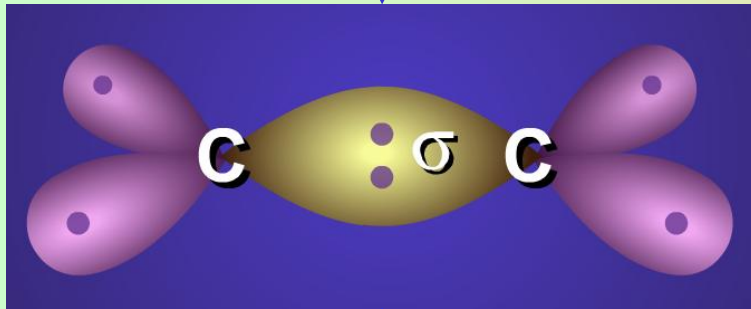
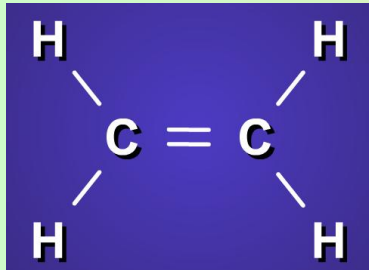


Sigma loturak

Lotura simple guztiak, abiapuntua orbital hibridoak, s edo p orbitalak izanik elkarsartzen direnak (σ) loturak dira.



Sigma eta Pi Loturak



Pi loturak

Pi loturak ematen dira p orbitalak konbinatzen direnean eta lotura-mota honek atomo baten bestearekiko errotazioa ekiditzen du; errotazioak pi lotura apurtuko luke.

