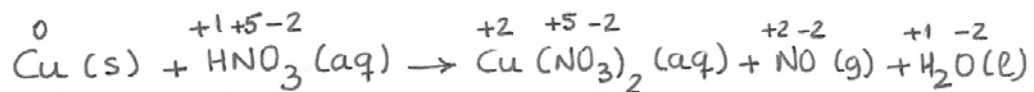


Uztaila - 2006

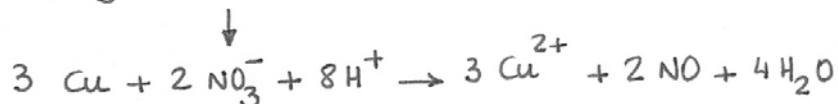
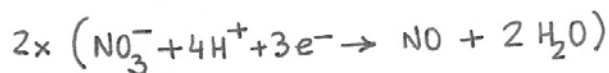
(A2) (a) Oxidazio-egoerak:



oxidatua ... Cu ...  $\text{Cu} \rightarrow \text{Cu}^{2+}$

erreduzitua: N ...  $\text{NO}_3^- \rightarrow \text{NO}$

(b) Ekuazio doituak



Erreaktibo mugatzailearen determinazioa

$$n(\text{HNO}_3) = 3 \frac{\text{mol}}{\text{L}} \times 0.1 \text{L} = 0.3 \text{ mol HNO}_3$$

10 g Cu behar duen HNO<sub>3</sub>:

$$n(\text{HNO}_3) = 10 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.5 \text{ g Cu}} \times \frac{8 \text{ mol HNO}_3}{3 \text{ mol Cu}} = 0.42 \text{ mol HNO}_3$$

behar du  
HNO<sub>3</sub> mugatzailea

Sortzen den Cu(NO<sub>3</sub>)<sub>2</sub> masa, gehienez:

$$m(\text{Cu}(\text{NO}_3)_2) = 0.3 \text{ mol HNO}_3 \times \frac{3 \text{ mol Cu}(\text{NO}_3)_2}{8 \text{ mol HNO}_3} \times \frac{187.5 \text{ g}}{1 \text{ mol Cu}(\text{NO}_3)_2}$$

$$M_m(\text{Cu}(\text{NO}_3)_2) = (1 \times 63.5) + (2 \times 14) + (6 \times 16) = 187.5 \text{ g/mol}$$

$$\boxed{m = 21.1 \text{ g Cu}(\text{NO}_3)_2}$$