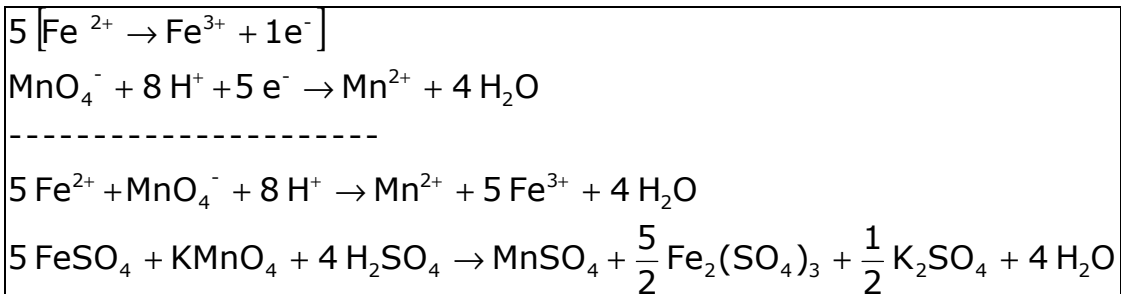
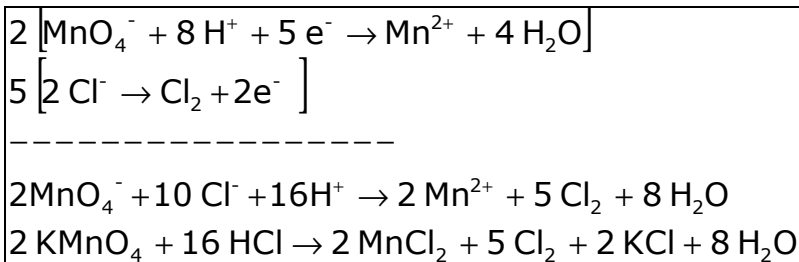
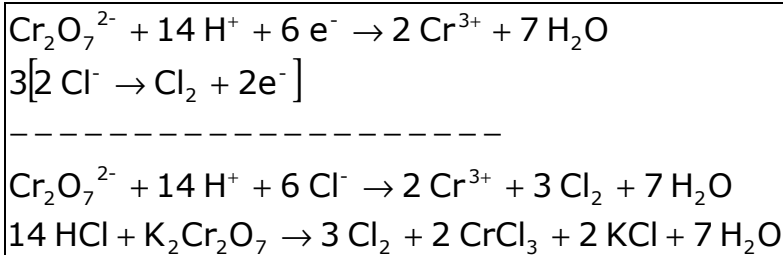
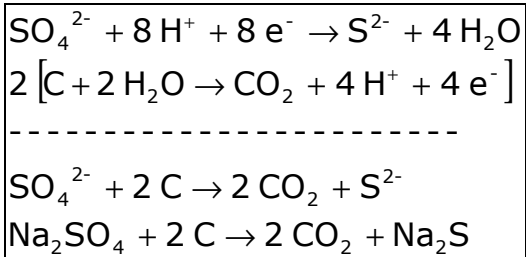
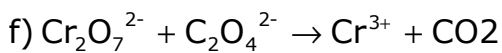
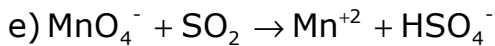
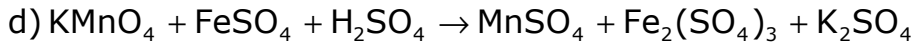
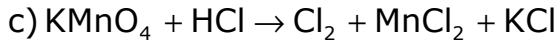
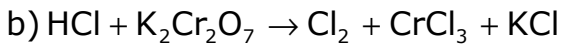
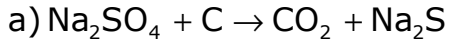
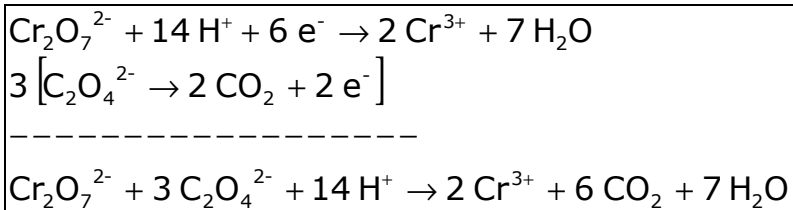
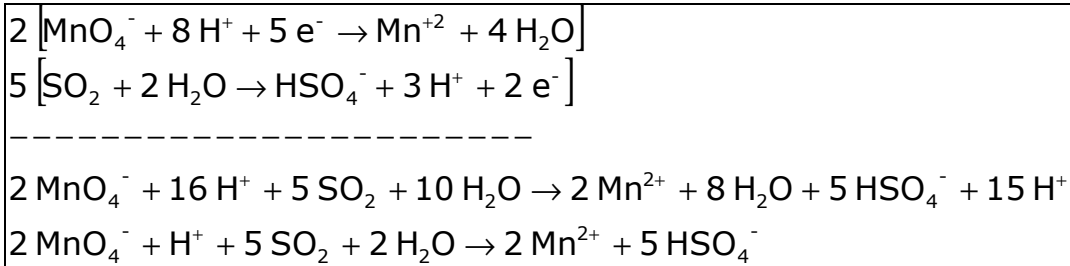


Oxidazio-erredukziozko ekuazioen doikuntza: ariketak

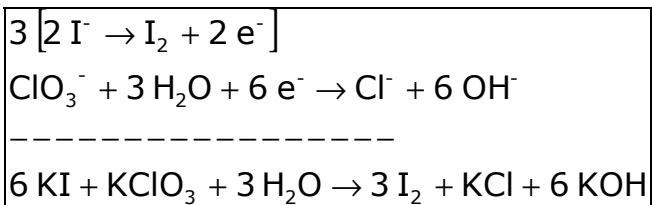
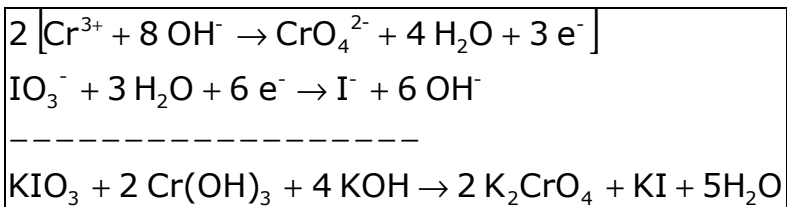
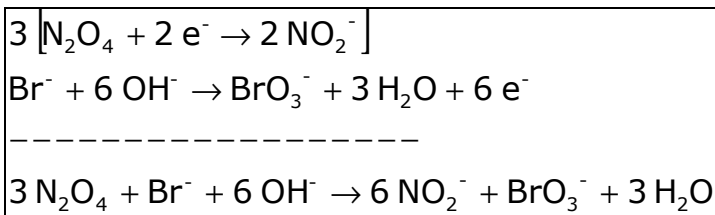
1. Ondoko ekuazioak ingurune azidoko redox erreakzioei dagozkie, doi itzazu:





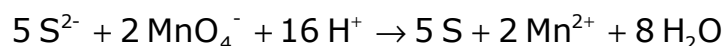
2. Ondoko ekuazioak ingurune basikoko redox erreakzioei dagozkie, doi itzazu:

- a) $\text{N}_2\text{O}_4 + \text{Br}^- \rightarrow \text{NO}_2^- + \text{BrO}_3^-$
 b) $\text{Cr}(\text{OH})_3 + \text{KIO}_3 \rightarrow \text{KI} + \text{K}_2\text{CrO}_4$
 c) $\text{KI} + \text{KClO}_3 \rightarrow \text{I}_2 + \text{KCl} + \text{KOH}$

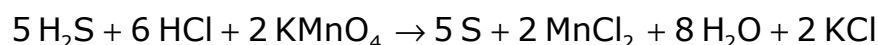


3. Potasio permanganatoak ingurune azidoan hidrogeno sulfuroa sufre elementaleraino, S, oxidatzeko gai da, eta permanganatoa manganeso (II) ioiraino pasatzen da. Oxidazio-erredukziozko erreakzioa doitu, forma molekularrean paratu, eta adierazi zein diren oxidatzailea, erreduktorea, oxidatzen den espeziea eta erreduzitzen den espeziea.

5 $[S^{2-} \rightarrow S + 2 e^-]$ S^{2-} oxidatzen da; erreduktorea
 2 $[MnO_4^- + 8 H^+ + 5 e^- \rightarrow Mn^{2+} + 4 H_2O]$ erreduzitzen da; MnO_4^- oxidatzailea



Ingurune azidoa HCl balitz:



4. Ioi-elektroi metodoa erabiliz, ondoko erreakzioa doitu: potasio dikromatoa + etanola + azido sulfurikoa → kromo (III) sulfatoa + azido etanoikoa + potasio sulfatoa + ura

