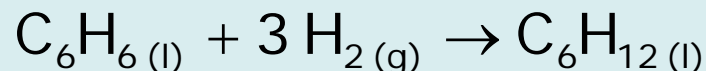
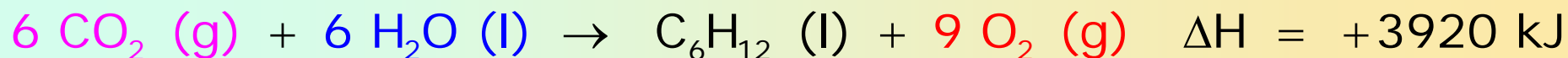
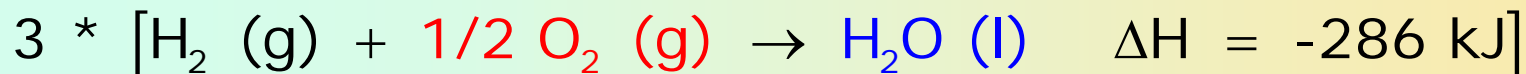


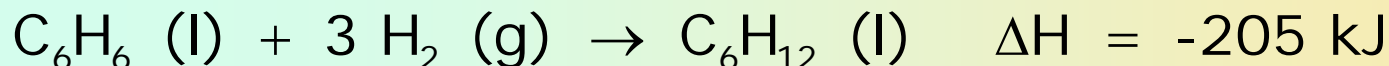
Entalpia-aldaketak: emaitzak

1. Kalkulatu bentzenoaren hidrogenazioaren entalpia-aldaketa, ziklohexanoa ematen duenean:



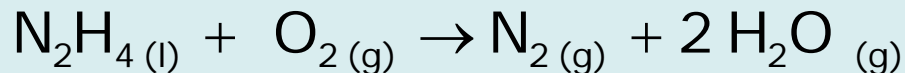
Konbustioaren entalpia-aldaketa estandarrek (kJ/mol-etan):
bentzenoa -3267, hidrogenoa -286, ziklohexanoa -3920





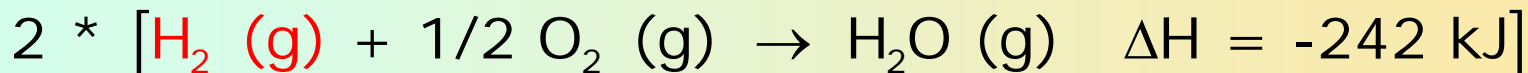
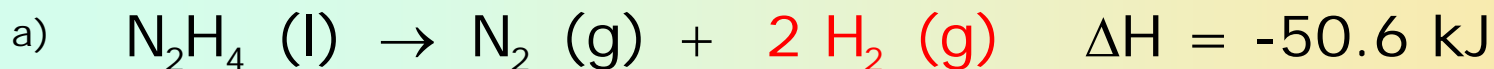
Entalpia-aldaketak: emaitzak

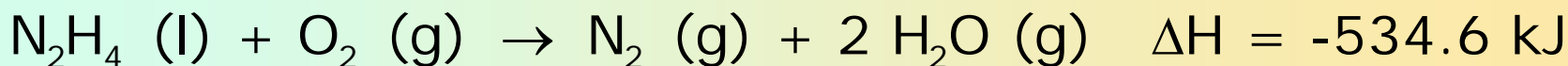
2. Hidrazina, N_2H_4 (l), espazio-untzien erregai gisa erabiltzen da. Oxigenoan erretzen da nitrogeno eta ur-lurrina emateko:



a) Kalkulatu entalpia-aldaketa 1 mol hidrazina erretzen denean. Formazio-entalpia estandarrak (kJ/mol-etan): hidrazina +50,6; H_2O (g) -242

b) Zenbateko bero-trukaketa ematen da hidrazina tona bat guztiz erretzen denean? (H=1; N=14; 1 tona=1000 kg)





b)

$$M = (2 * 14) + (4 * 1) = 32 \text{ g/mol}$$

$$Q = \frac{-534.6 \text{ kJ}}{1 \text{ mol}} * \frac{1 \text{ mol}}{32 \text{ g}} * 10^6 \text{ g} = -1.67 * 10^7 \text{ kJ}$$

Entalpia-aldaketak: emaitzak

3. Kalkulatu bromoetanoaren formazio-entalpia, $C_2H_5Br_{(l)}$ ondoko datuak erabiliz

