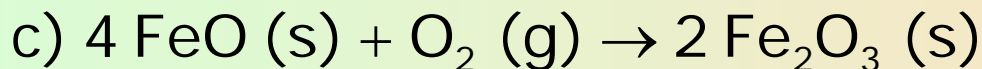
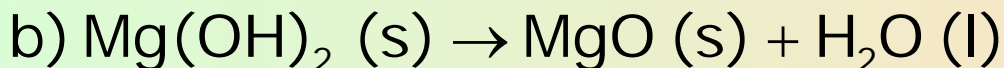
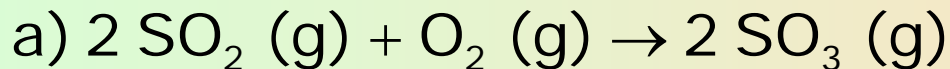


Ondoko erreakzioen entalpia-aldaketa estandarrak kalkulatu:



Formazio-entalpiak:

$$\Delta H_f^\circ [\text{SO}_2 (\text{g})] = -296.9 \text{ kJ/mol}; \quad \Delta H_f^\circ [\text{SO}_3 (\text{g})] = -395.2 \text{ kJ/mol};$$

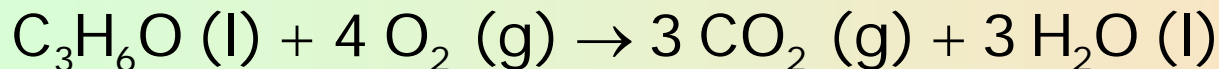
$$\Delta H_f^\circ [\text{Mg}(\text{OH})_2 (\text{s})] = -924.7 \text{ kJ/mol}; \quad \Delta H_f^\circ [\text{MgO} (\text{s})] = -601.8 \text{ kJ/mol}$$

$$\Delta H_f^\circ [\text{FeO} (\text{s})] = -271.9 \text{ kJ/mol}; \quad \Delta H_f^\circ [\text{Fe}_2\text{O}_3 (\text{s})] = -822.2 \text{ kJ/mol}$$

$$\Delta H_f^\circ [\text{H}_2\text{O} (\text{l})] = -285.8 \text{ kJ/mol}$$

Formazio-entalpiak: ariketak

1 mol azetona, C_3H_6O , guztiz erretzen denean 1790 kJ askatzen ditu:



Kalkulatu azetonaren formazio-entalpia.

Entalpia-aldaketak: $\Delta H_f^\circ [CO_2 (g)] = -393.5 \text{ kJ/mol};$

$\Delta H_f^\circ [H_2O (l)] = -285.8 \text{ kJ/mol};$

Kalkulatu diborano gaseosoaren, B_2H_6 , formazio-entalpia, ondoko informazio termokimikoa kontuan hartuz:

