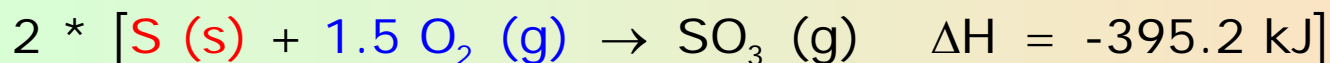
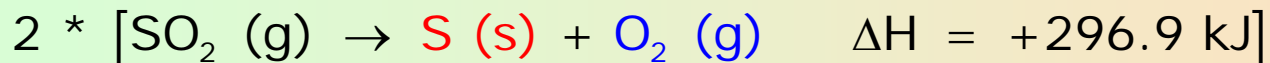
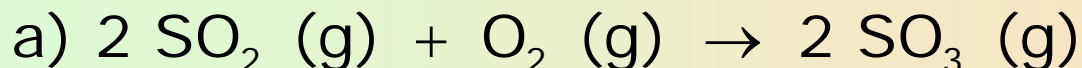
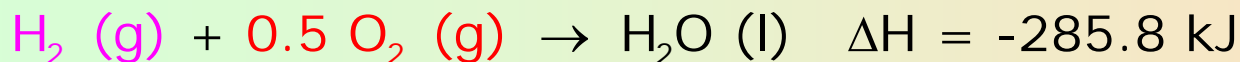
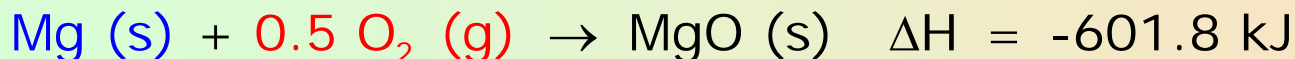
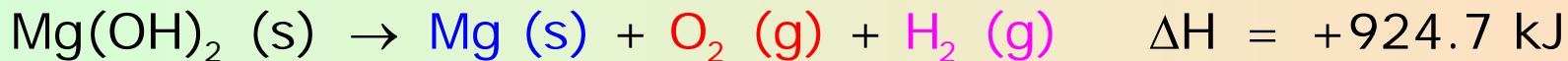
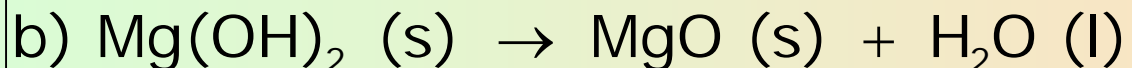
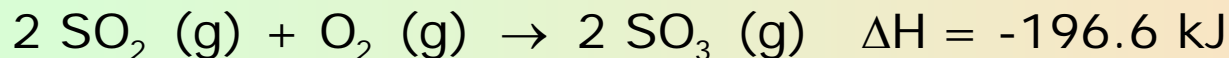
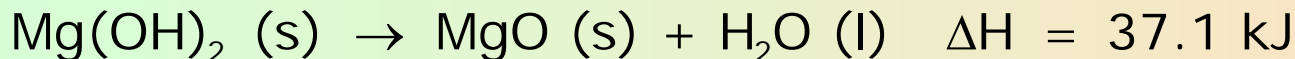


Formazio-entalpiak: emaitzak

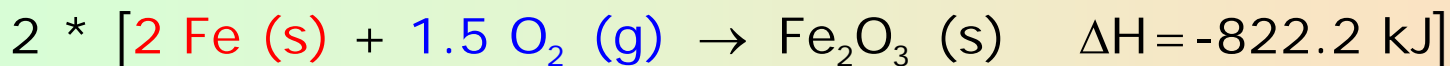
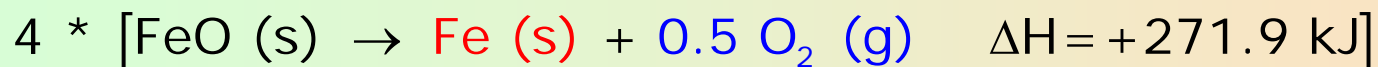
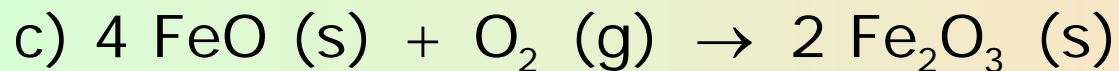
Ondoko erreakzioen entalpia-aldaketa estandarrak kalkulatu:

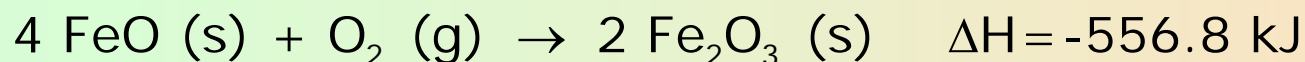






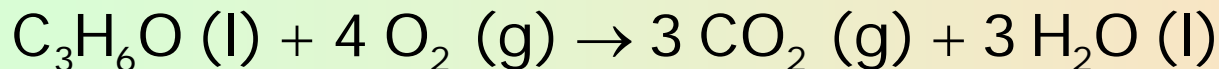
Ondoko erreakzioen entalpia-aldaketa estandarrak kalkulatu:





Formazio-entalpiak: emaitzak

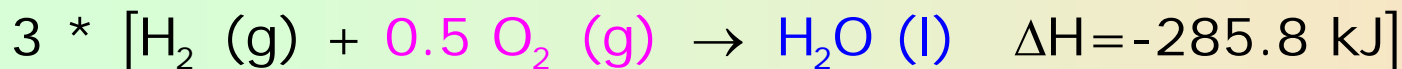
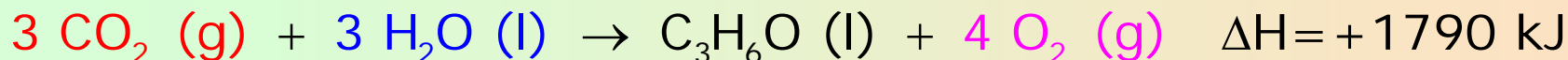
1 mol azetona, C_3H_6O , guztiz erretzen denean 1790 kJ askatzen ditu:

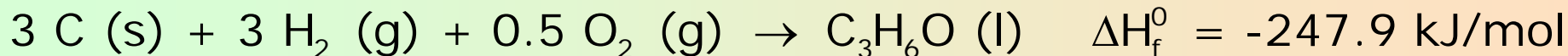


Kalkulatu azetonaren formazio-entalpia.

Entalpia-aldaketak: $\Delta H_f^0 [CO_2 (g)] = -393.5 \text{ kJ/mol};$

$\Delta H_f^0 [H_2O (l)] = -285.8 \text{ kJ/mol};$





Formazio-entalpiak: emaitzak

Kalkulatu diborano gaseosoaren, B_2H_6 , formazio-entalpia, ondoko informazio termokimikoa kontuan hartuz:

